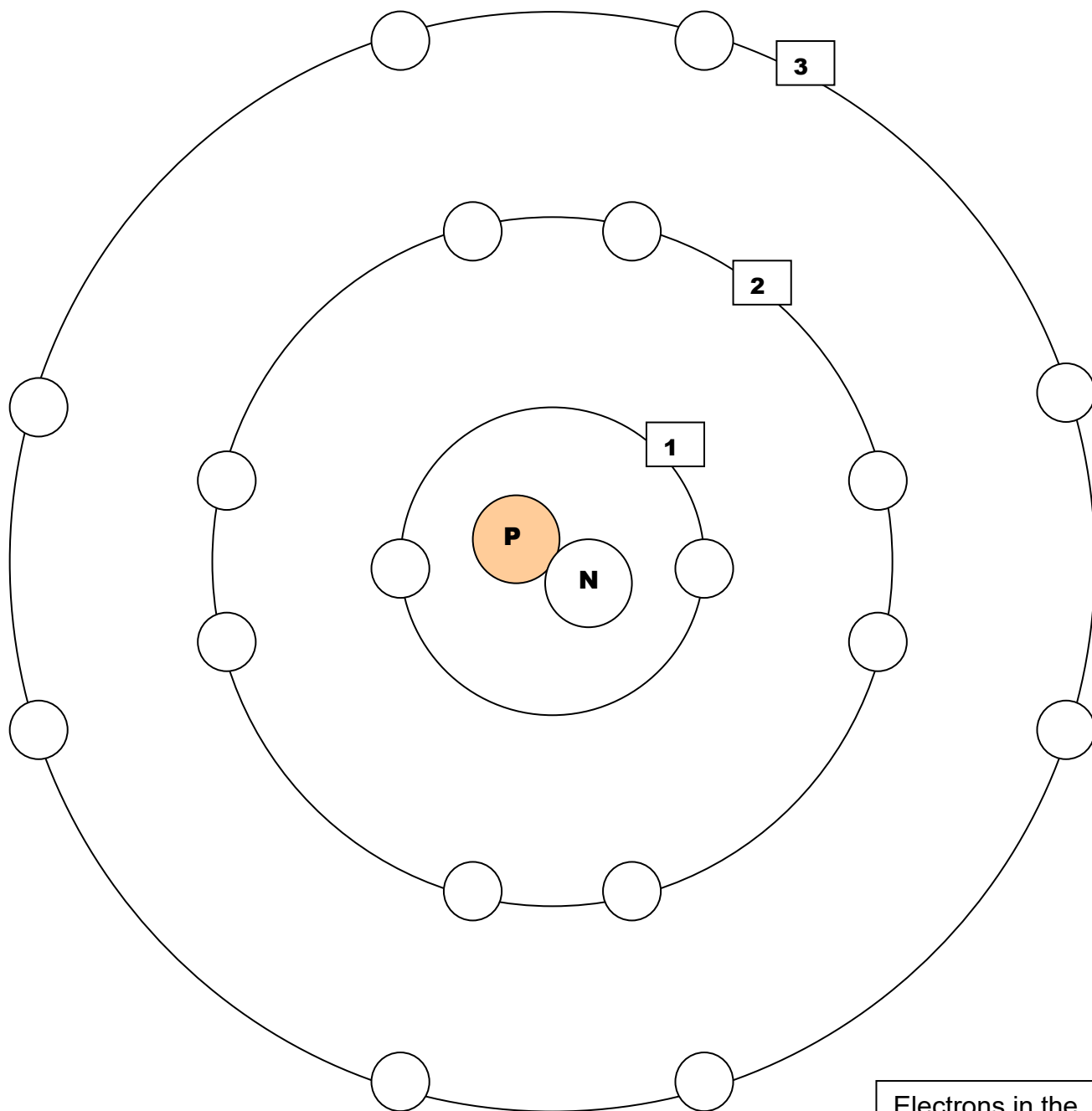


Name: \_\_\_\_\_

## Atoms and Molecules

# ORBITALS



**Element = \_\_\_\_\_**

Atomic Number = \_\_\_\_\_

Number of Protons = Atomic Number = \_\_\_\_\_

Atomic Mass (rounded to nearest whole number) = \_\_\_\_\_

Number of Neutrons = Atomic Mass – Atomic Number = \_\_\_\_\_

Number of Electrons = Number of Protons = \_\_\_\_\_

Period = \_\_\_\_\_ = Number of Orbitals

Group = \_\_\_\_\_ = Number of electrons in outermost orbital

Electrons in the outermost orbital [3 in this case] are called valence electrons.